Centurion
UNIVERSITY
Shaping Lives...
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B.Sc. (Chemistry) Syllabus

(Three Years Programme)

School of Applied Sciences

Centurion University of Technology & Management

2024-25

Department of Chemistry

Sl No	Course Code	Course Name	Credit	Type [T+Pr+Pj]
1	CUTM1469	Atomic Structure and Chemical bonding	6	3+2+1
2	CUTM1470	States of matter and ionic equilibrium	6	3+2+1
3	CUTM1471	Basics and Hydrocarbons	6	3+2+1
4	CUTM1472	Chemical Thermodynamics and its application	6	3+2+1
5	CUTM1473	S- and P-block elements	6	3+2+1
6	CUTM1474	Oxygen Containing Functional Groups	6	3+2+1
7	CUTM1475	Phase Equilibria & Chemical Kinetics	6	3+2+1
8	CUTM1476	Coordination Chemistry	6	3+2+1
9	CUTM1477	Heterocyclic Chemistry	6	3+2+1
10	CUTM1478	Electrochemistry	6	3+2+1
11	CUTM1479	Bio-molecules	6	3+2+1
12	CUTM1480	Quantum Chemistry & Spectroscopy	6	3+2+1
13	CUTM1481	Organometallic chemistry	6	3+2+1
14	CUTM1482	Spectroscopy	6	3+2+1

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1469	Atomic Structure and Chemical Bonding	3-2-1	

Objective

- To develop basic concepts of theories of atoms and molecules.
- To understand the most common structure and hybridization observed for different compounds.
- To use periodic trends in order to understand the chemistry of alkali metals, alkaline earth metals, halogens, transition as well as inner transition elements.

Course outcome

After completion of the course, students will be able to,

COs	Course outcomes
CO1	Acquire knowledge in the structure, bonding, and hybridisation of different compounds, their periodic properties and the quantum mechanical aspect.
CO2	Solve problems of different wave functions and wave equations.
CO3	Improved experimental and theoretical knowledge by analyzing the data.
CO4	Develop problem solving aptitude related to behavior of different compounds
CO5	Develop innovative ideas to apply in journal publications.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	2	2	2			2			2	2	3	2	3
CO2	3	3	3	2	2			2			2	2	3	2	3
CO3	3	3	3	2	2			2			2	2	3	2	3
CO4	-	3	2	2	2			2			2	2	3	2	3
CO5	-	3	2	2	2			2			2	2	3	2	3

***High-3, Medium-2, Low-1**

Course content

Module I

Atomic structure-I (10h)

Bohr's theory, its limitations and atomic spectrum of hydrogen atom, Wave mechanics: de Broglie equation, Heisenberg's Uncertainty Principle and its significance, Schrodinger's wave equation, significance of Ψ and Ψ^2.

Assignment-1: Photoelectric Effect

Assignment-2: Sommerfeld Extension of the Bohr Theory

Module II

Atomic structure-II (12h)

Quantum numbers and their significance, Normalized and orthogonal wave functions. Sign of wave functions. Radial and angular wave functions for hydrogen atom. Radial and angular distribution curves. Shapes of s, p, d and f orbital's. Contour boundary and probability diagrams. Pauli's Exclusion Principle, Hund's rule of maximum multiplicity, Aufbau's principle and its limitations, Variation of orbital energy with atomic number.

Assignment-3: Atomic term symbol

Module III

Periodic Properties-I (9h)

s, p, d, f block elements, the long form of periodic table. Detailed discussion of the following properties of the elements with reference to s and p-block, Effective nuclear charge, shielding or screening effect, Slater rules, variation of effective nuclear charge in periodic table, Atomic radii (van der Waals), Ionic and crystal radii, covalent radii (octahedral and tetrahedral), Ionization enthalpy, Successive ionization enthalpies and factors affecting ionization energy, Applications of ionization enthalpy

Assignment-4: Application of effective nuclear charge

Assignment-5: Shielding effect of inner-shell electrons on the valence shell electron

Module IV

Periodic Properties-II (10h)

Electron gain enthalpy, trends of electron gain enthalpy. Electronegativity, Pauling's/ Mulliken's electronegativity scales. Variation of electronegativity with bond order, partial charge, hybridization, group electronegativity.

Assignment-6: Polarizing powder and Polarizability

Practice-1: Estimation of free alkali present in different soaps/detergents (Vlab)

Practice-2: Estimation of Barium (Vlab).

Practice-3: Estimation of Fe(II) and oxalic acid using standardized KMnO4 solution.

Module V

Chemical Bonding-I(Ionic Bonding) (7h)

Ionic bond: General characteristics, types of ions, size effects, radius ratio rule and its limitations, Packing of ions in crystals, Born-Lande equation Madelung constant, Born-Haber cycle and its application, solvation energy.

Assignment -7: Radius ratio of different cubic crystal structures

Practice-4: Determine the solubilities of solid compounds in water and in ethanol

Module VI

Chemical Bonding-II (Covalent Bonding-I) (9h)

Covalent bond: Lewis structure, Valence Bond theory (Heitler-London approach), Energetic of hybridization, equivalent and non-equivalent hybrid orbitals, Resonance and resonance energy, Molecular orbital theory. Molecular orbital diagrams of diatomic and simple polyatomic molecules, Formal charge, Valence shell electron pair repulsion theory (VSEPR).

Assignment-8: Wave mechanical treatment of covalent bond

Practice-5: Compare the melting points of various compounds.

Module VII

Chemical Bonding-III (Covalent Bonding-II) (9h)

Covalent character in ionic compounds, polarizing power and polarisability, Fajan's rules and consequences of polarization, Ionic character in covalent compounds: Bond moment and dipole moment, Percentage ionic character from dipole moment and electronegativity difference

Assignment-9: Molecular orbital's for homonuclear diatomic molecules

Assignment-10: Term symbols for a diatomic molecule

Practice-6: Determine the conductivity of water solutions of the soluble solids

Book References:

- Lee, J.D. Concise Inorganic Chemistry, ELBS, 1991.
- Douglas, B.E. and Mc Daniel, D.H., Concepts & Models of Inorganic Chemistry, Oxford, 1970
- Atkins, P.W. & Paula, J. Physical Chemistry, Oxford Press, 2006.
- Day, M.C. and Selbin, J. Theoretical Inorganic Chemistry, ACS Publications 1962.

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1470	States of Matter and Ionic Equilibrium	3-2-1	

Objective

- To develop basic and advance concepts regarding the three states of matter.
- To derive the expressions for determining the physical properties of gases, liquids and solids.

Course outcome:

After completion of the course, students will be able to

COs	Course outcomes
CO1	Recall on the thermodynamic, colligative, and electrical properties of sample
CO2	Describe on the handling of pH meter and it's usage in various sample analysis
CO3	Demonstrate the concept of ionization of both weak and strong electrolytes
CO4	Analyze the behavior of different types weak acids and bases in solution phase
CO5	Develop innovative idea for synthesis of product and journal publications.

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	2	2	2	3	2	2	3	1	2	1	3	2	1
CO2	3	2	3	1	2	1	2	3	1	1	3	2	3	1	2
CO3	2	3	1	2	2	2	1	1	2	3	2	1	2	2	2
CO4	2	2	3	2	1	1	2	1	1	3	2	2	1	3	2

CO5	2	2	1	3	1	2	3	2	3	2	2	3	2	2	3

***High-3, Medium-2, Low-1**

Course content

Module I

Gaseous state-I (7h)

Deviation of real gases from ideal behaviour, compressibility factor, causes of deviation. Van der Waals equation of state for real gases, Boyle temperature (derivation not required), Critical phenomena, critical constants and their calculation from van der Waals equation.

Assignment-1: Empirical Gas Laws

Assignment-2: Ideal gas equation

Module II

Gaseous state-II (8h)

Maxwell Boltzmann distribution laws of molecular velocities, Molecular energies (graphic representation – derivation not required) and their importance, Temperature dependence of these distributions. Most probable, average and root mean square velocities, Collision cross section, collision number, collision frequency, collision diameter and mean free path of molecules. Viscosity of gases and effect of temperature and pressure on coefficient of viscosity (qualitative treatment only)

Assignment-3: Degrees of freedom of the gaseous molecule

Assignment-4: Liquefaction of gases

Module III

Liquid state (13h)

Qualitative treatment of the structure of the liquid state, Physical properties of liquids; vapour pressure, surface tension and coefficient of viscosity, and their determination, Effect of addition of various solutes on surface tension and viscosity. Temperature variation of viscosity of liquids and comparison with that of gases

Assignment-5: Thermodynamic and intermolecular forces in solution

Practice-1: Determination of surface tension

Practice-2: Determination of Viscosity of Organic Solvents

Practice-3: Determination of co-efficient of viscosity of an unknown aqueous solution

Practice-4: Study the variation of surface tension of detergent solutions with concentration

Module IV

Solid state (7h)

Nature of the solid state, law of constancy of interfacial angles, law of rational indices, Miller indices, Elementary ideas of symmetry, symmetry elements and symmetry operations, qualitative idea of point and space groups, seven crystal systems and fourteen Bravais lattices. X-ray diffraction, Bragg's law, a simple account of rotating crystal method and powder pattern method, Analysis of powder diffraction patterns of NaCl, CsCl and KCl

Assignment-6: Calculating the intensity of diffraction using the structure factor equation

Module V

Ionic equilibria-I (10h)

Strong, moderate and weak electrolytes, Degree of ionization, factors affecting degree of ionization, ionization constant and ionic product of water. Ionization of weak acids and bases, pH scale, common ions effect, dissociation constants of mono and diprotic acids

Assignment -7: Strength of acids and bases

Practice-5: Determination of the pH scale by the method of successive dilutions

Practice-6: Study the effect of addition of HCl/NaOH on pH to the solutions of acetic acid, sodium acetate and their mixtures.

Module VI

Ionic equilibria-II (10h)

Degree of ionization, factors affecting degree of ionization, Ionization constant and ionic product of water, Buffer solutions; derivation of Henderson equation and its applications, Solubility and solubility product of sparingly soluble salts – applications of solubility product principle

Assignment-8: Mixture of two weak acids

Assignment-9: Ionization of polyprotic acids

Module VII

Ionic equilibria-III (11h)

Qualitative treatment of acid – base titration curves, Calculation of pH at various stages, Theory of acid–base indicators, Selection of indicators and their limitations

Assignment-10: Common acid-base indicators and their properties

Book References:

Theory:

Atkins, P.W.; Paula, J.de. (2014),Atkin's Physical Chemistry Ed., 10th Edition, Oxford University Press.
Ball, D. W. (2017),Physical Chemistry, 2nd Edition,Cengage Learning, India.
Castellan, G. W. (2004),Physical Chemistry, 4th Edition, Narosa.
Kapoor, K.L. (2015),A Textbook of Physical Chemistry, Vol 1, 6th Edition, McGraw Hill Education.
Practical:
Khosla, B.D.; Garg, V.C.; Gulati, A. (2015),Senior Practical Physical Chemistry, R. Chand & Co, New Delhi.
Kapoor, K.L. (2019),A Textbook of Physical Chemistry, Vol.7, 1st Edition, McGraw Hill Education.
Garland, C. W.; Nibler, J. W.; Shoemaker, D. P.(2003),Experiments in Physical Chemistry, 8th Edition, McGraw-Hill, New York.

Additional Resources:

Moore, W.J. (1972),**Physical Chemistry**, 5th Edition,Longmans Green & Co. Ltd.

Glasstone, S. (1948), **Textbook of Physical Chemistry**, D. Van Nostrand company

Note: 1 credit theory=10 hrs lecture, 1 credit practice/project=12.5 hrs lab/workshop/field work in a semester

Course Outline

Code	Course Title	T-P-P (Credit)	Prerequisite
CUTM1471	Basics of Hydrocarbons	3-2-1	

Objective

- The general concept of this course is to train students the fundamental laboratory skills like extraction, purification and separation techniques with some simple organic preparations.
- This helps students to gain experience to predict the functional group transformations, simple reaction mechanisms, and the synthesis of organic molecules by multi-step synthesis strategies

Course outcome

COs	Course outcomes
CO1	Able to gain knowledge on various type of functional group transformations and its applications in multi-step synthesis of organic drug molecules.
CO2	Able to gain skill on interpretation of the behavior of different organic compounds.
CO3	Investigate the reaction pathways pertaining to hydrocarbons
CO4	Able to get skill on extraction, purification and separation techniques and can be applied to organic synthesis.
CO5	Develop innovation for idea for synthesis of o**rganic** product and **also for** journal publications.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	3	3	2	3					1	2	1	1	3	3
CO2	3	2	2	2	3					3	2	3	3	2	3
CO3	2	3	1	2	3					1	3	2	2	2	3

CO4	2	2	2	1	3					2	1	3	3	1	3
CO5	2	3	3	2	2					2	2	1	2	2	3

***High-3, Medium-2, Low-1**

Course content

Module -I: Basics of Organic Chemistry(5hrs)

Organic Compounds: Classification, and Nomenclature, Hybridization, Shapes of molecules, Influence of hybridization on bond properties.Electronic Displacements: Inductive, electromeric, resonance and mesomeric effects, hyperconjugation and their applications. Dipole moment; Organic acids and bases; their relative strength.

Assignment-1 : Organic Compounds, Electronic Displacements.

Assignment-2 : Electronic Displacements, Organic acids and bases; their relative strength

Module – II : Reaction Intermediate(4hrs)
Homolytic and Heterolytic fission with suitable examples. Curly arrow rules, formal charges; Electrophiles and Nucleophiles; Nucleophlicity and basicity. Types, shape and their relative stability of Carbocations, Carbanions, Free radicals and Carbenes. Introduction to types of organic reactions and their mechanism: Addition, Elimination and Substitution reactions.

Assignment-3 : Reaction Intermediate

Assignment-4 : Introduction to types of organic reactions and their mechanism.

Module – III : Chemistry of Aliphatic Hydrocarbons(2hrs)

Chemistry of alkanes: Formation of alkanes, Wurtz Reaction, Wurtz-Fittig Reactions.Free radical substitutions: Halogenation -relative reactivity and selectivity.

Practice-1 : Determination of the melting points of above compounds and unknown organic compounds

Practice-2 : Effect of impurities on the melting point - mixed melting point of two unknown organic compounds.

Module – IV : **Carbon-Carbon pi bonds(8hrs)**

Formation of alkenes and alkynes by elimination reactions. Mechanism of E1, E2, E1cb reactions. Saytzeff and Hofmann eliminations. Reactions of alkenes: Electrophilic additions their mechanisms (Markownikoff/ Anti-Markownikoff addition). Mechanism of oxymercuration-demercuration, hydroboration-oxidation, ozonolysis, reduction (catalytic and chemical), syn and anti-hydroxylation (oxidation) 1,2-and 1,4-addition reactions in conjugated dienes and, Diels-Alder reaction; Allylic and benzylic bromination and mechanism, e.g. propene, 1-butene, toluene, ethyl benzene. Reactions of alkynes: Acidity, Electrophilic and Nucleophilic additions. Hydration to form carbonyl compounds, Alkylation of terminal alkynes.

Practice-3 : Separation of Amino acids by TLC

Assignment-5 : Elimination and Substitution reaction

Assignment-6 : Alkene,Reaction of Alkene

Module – V : Aromatic Hydrocarbons(3hrs)

Aromaticity: Hückel's rule, aromatic character of arenes, cyclic carbocations/carbanions and heterocyclic compounds with suitable examples. Electrophilic aromatic substitution: halogenation, nitration, sulphonation. Electrophilic aromatic substitution: Friedel-Craft's alkylation/acylation with their mechanism. Directing effects of the groups.

Practice-4 : Detection of N,S,Cl,Br and I in organic compound.

Assignment-7 : Reactions of alkynes, Aromaticity, Electrophilic aromatic substitution.

Module - VI : Cycloalkanes and Conformational Analysis: (2hrs)

Types of cycloalkanes and their relative stability, Baeyer strain theory. Conformation analysis of alkanes. Relative stability: Energy diagrams of cyclohexane: Chair, Boat and Twist boat forms; Relative stability with energy diagrams.

Assignment-8 : Cycloalkanes

Module – VII : Stereochemistry(6hrs)

Fischer Projection, Newmann and Sawhorse Projection formulae and their interconversions. Geometrical isomerism: cis–trans and, syn-anti isomerism E/Z notations with C.I.P rules. Optical Isomerism: Optical Activity, Specific Rotation, Chirality/Asymmetry, Enantiomers, Molecules with two or more chiral-centres. Distereoisomers, meso structures, racemic mixture and resolution. Relative and absolute configuration: D/L and R/S designations.

Practice-5 : Preparation of cyclohexane from cyclohexanol.

Assignment-9 : Conformation analysis of alkanes, Relative stability, Fischer Projection, Newmann and Sawhorse Projection formulae and their inter conversions .

Assignment-10 : Stereochemistry.

Text Books:

1. University Chemistry, Vol. II, U.N Ojha and K.K Ojha, Himalaya Publishing House

2.Modern College Chemistry, Y.R Sharma and K.D Sharma, Kalyani Publishers

Reference Books:

1.Morrison, Boyd and Bhattacharjee, Organic Chemistry, 7th Edition, Pearson.
2.Kalsi, P. S. Stereochemistry Conformation and Mechanism, New Age International, 2005.
3.Eliel, E. L. & amp; Wilen, S. H. Stereochemistry of Organic Compounds, Wiley: London, 1994.

Course Outline.

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1472	Chemical Thermodynamics and its Applications	**(3-2-1)**	

Objective

- To develop basic and advanced concepts regarding fundamentals thermodynamics and its related phenomena.
- To explore the concept of chemical equilibrium in attaining high products by controlling various factors.
- To identify the concept of ideal and non-ideal solution along with various colligative properties.

Course outcome

After completion of the course, students will be able to,

COs	Course outcomes
CO1	Gain Knowledge on different thermodynamic laws, Chemical Equilibria and Colligative properties.
CO2	Describe different derivations for thermodynamic equations, chemical equilibriums and their relationship among each other.
CO3	Develop problem solving skills related to thermodyanmics.
CO4	Correlate thermodynamic derivation of different equilibrium constants.
CO5	Develop innovative ideas to apply in scientific research for journal publications.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	2	3	3					2				1	2		1
CO2	2	3	2	2		2			1				2	2	1
CO3	2	3	3	2		2				2			3	3	
CO4	3	3	3	2		2					3		3	3	
CO5	3	3	3	2	2			2					2		1

***High-3, Medium-2, Low-1**

Course content

Module I : Introduction to Thermodynamics (6 hrs)

Intensive and extensive variables; state and path functions; isolated, closed and open systems Zeroth law of thermodynamics, First law: Concept of heat (q), work (w), internal energy (U), and statement of first law, Enthalpy (H), relation between heat capacities, Calculations of q, w, U and H for reversible condition, Calculations of q, w, U and H for irreversible and free expansion of gases (ideal and van der Waals) under isothermal and adiabatic conditions

Module II : Thermochemistry (12 hrs)

Heats of reactions: standard states; Enthalpy of formation of molecules and ions and enthalpy of combustion and its applications, Effect of temperature (Kirchhoff's equations) and pressure on enthalpy of reactions, Adiabatic flame temperature, explosion temperature, Calculation of bond energy, bond dissociation energy and resonance energy from thermochemical data

Practice 1: Determination of heat capacity of a calorimeter for different volumes using change of enthalpy data of a known system (method of back calculation of heat capacity of calorimeter from known enthalpy of solution or enthalpy of neutralization).

Practice 2: Determination of heat capacity of the calorimeter and enthalpy of neutralization of hydrochloric acid with sodium hydroxide.

Practice 3: Determination of heat capacity of the calorimeter and integral enthalpy (endothermic and exothermic) solution of salts.

Practice 4: Determination of enthalpy of hydration of copper sulphate

Module III : Chemical Thermodynamics I (9 hrs)

Thermodynamic scale of temperature, Statement of the second law of thermodynamics; molecular and statistical interpretation of entropy, Calculation of entropy change for reversible and irreversible processes, Third Law: Statement of third law, concept of residual entropy , calculation of absolute entropy of molecules, Free Energy Functions: Gibbs and Helmholtz energy; variation of S, G, A with T, V, P; Free energy change and spontaneity, Relation between Joule-Thomson coefficient and other thermodynamic parameters; inversion temperature, Thermodynamic equation of state, Gibbs-Helmholtz equation, Maxwell relations

Module IV : Chemical Thermodynamics and Open System (5 hrs)

Partial molar quantities, dependence of thermodynamic parameters on composition, Gibbs Duhem equation, chemical potential of ideal mixtures, change in thermodynamic functions in mixing of ideal gases, Partial molar quantities

Module V : Chemical Equilibrium: I (10 hrs)

Criteria of thermodynamic equilibrium, Degree of advancement of reaction, chemical equilibria in ideal gases, Thermodynamic derivation of relation between Gibbs free energy of reaction and reaction quotient, Concept of fugacity, Equilibrium constants and their quantitative dependence on temperature, pressure and concentration, Coupling of exoergic and endoergic reactions, Free energy of mixing and spontaneity

Practice 5: To study the shift of equilibrium between ferric ions and thiocyanate ions by increasing the concentration of either of them

Module VI : Chemical Equilibrium: II (8 hrs)

Thermodynamic derivation of relations between the various equilibrium constants Kp, Kc and Kx. Le Chatelier principle (quantitative treatment), Equilibrium between ideal gases and a pure condensed phase, Introduction to dilute Solutions, Introduction to Colligative Properties

Module VII: Colligative Properties (16 hrs)

Excess thermodynamic functions, Thermodynamic derivation using chemical potential to derive relations between the relative lowering of vapour pressure and amount of solute, Raoult's and Henry's Laws and their applications, Thermodynamic derivation using chemical potential to derive relations between the elevation of boiling point and amount of solute, Thermodynamic derivation using chemical potential to derive relations between Depression of freezing point and amount of solute, Thermodynamic derivation using chemical potential to derive relations between the osmotic pressure and amount of solute, Applications in calculating molar masses of normal, dissociated and associated solutes in solution

Practice 6: To find the molar mass through freezing point depression of a solution

Practice 7: To find the elevation of boiling point of a solution

Practice 8: To find the molar mass through elevation of boiling point of a solution

Text Books:

1. Atkins P. and De Paula, J. *Physical Chemistry* Tenth Ed., OUP, 2014.

2. Castellan, G. W. *Physical Chemistry 4th Ed.,* Narosa, 2004.

Reference Books:

1. Engel, T. and Reid, P. *Physical Chemistry 3rd Ed.,* Prentice Hall, 2012

2. .McQuarrie, D. A. and Simon, J. D. *Molecular Thermodynamics* Viva Books, 2004.

3. Roy, B. N. Fundamentals of Classical and Statistical Thermodynamics Wiley, 2001

4. Levine, I .N. *Physical Chemistry* 6th Ed. Tata Mc Graw Hill, 2010.

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1473	**S and P block elements**	**3-2-1**	

Objective

- To identify the common physical properties of metals and non- metals and explain how their uses relate to these properties.
- To explore in depth specialized areas of chemistry of materials, including ores, metals, acids and bases and to understand how metals are extracted from their ores.
- To understand the trends in properties and reactivity of the s, p-block elements and noble gases.
- To become familiar with some of the roles of inorganic polymer and its applications in day to day life

Course outcome

After Completion of the course students will be able to

COs	Course outcomes
CO1	Understand the fundamental aspects of alkali and alkaline earth metal including trends in oxidation states, periodic properties and complex formation tendency
CO2	Explain the role of metal in metallurgical chemical reaction
CO3	Demonstrate different methods for synthesizing inorganic materials and principles of inorganic polymer synthesis
CO4	Describe the role of alkali and alkaline earth metals in biological system
CO5	Experiment the inorganic polymer-based material and interpret the data.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4													3	2	3
CO5													3	2	3

***High-3, Medium-2, Low-1**

Course content

Module I (8 Hours)

General Principles of Metallurgy:

Chief modes of occurrence of metals based on standard electrode potentials, Ellingham diagrams for reduction of metal oxides using carbon and carbon monoxide as reducing agent, Electrolytic Reduction, Hydrometallurgy, Methods of purification of metals: Electrolytic Kroll process, Parting process, Mond's process, Zone refining.

Assignment-1

Practice-1

Gravimetric Estimation of Barium

Practice-2

Crystalisation of copper sulphate

Module II

Acids and Bases(7 Hours)

Bronsted-Lowry concept of acid-base reactions, solvated proton, relative strength of acids, types of acid-base reactions, levelling solvents, Lewis acid-base concept, Classification of Lewis acids, Hard and Soft Acids and Bases (HSAB) Application of HSAB principle.

Assignment-2

Assignment-3

Practice-3

Preparation of Manganese (III) phosphate, $MnPO_4.H_2O$.

Practice-4

Preparation of schiff base ligands

Module III(8 Hours)

Chemistry of s and p Block Elements-1:

Inert pair effect, Relative stability of different oxidation states, diagonal relationship, anomalous behaviour of first member of each group, Allotropy and catenation, Complex formation tendency of s and p block elements.

Assignment-4

Practice-5

Estimation of Cu(II) and K2Cr2O7 using sodium thiosulphate solution iodometrically

Module IV(9 Hours)

Chemistry of s and p Block Elements-2:

Hydrides and their classification ionic, covalent and interstitial, Basic beryllium acetate and nitrate, Study of the following compounds with emphasis on structure, bonding, preparation, properties and uses. Boric acid and borates, boron nitrides, borohydrides (diborane).

Assignment-6

Practice-6

Preparation of potash alum, K2SO4 Al2(SO4)3.24H2O

Module V(8 Hours)

Chemistry of s and p Block Elements-3:

Oxides and oxoacids of nitrogen, Phosphorus and chlorine, Peroxo acids of sulphur, interhalogen compounds, polyhalide ions, pseudohalogens and basic properties of halogens.

Assignment-7

Assignment-8

Practice-7

Estimation of available chlorine in bleaching powder

Practice-8

Gravimetric Estimation of Nickel

Module VI(8 Hours)

Noble Gases:

Occurrence and uses, rationalization of inertness of noble gases, Clathrates; preparation and properties of XeF2, XeF4 and XeF6, Nature of bonding in noble gas compounds (Valence bond treatment and MO treatment for XeF2), Molecular shapes of noble gas compounds (VSEPR theory).

Assignment-9

Practice-9

Preparation of Cuprous chloride, Cu2Cl2

Module VII(8 Hours)

Inorganic Polymers:

Types of inorganic polymers, comparison with organic polymers, synthesis, structural aspects and applications of silicones and siloxanes, Borazines, silicates, phosphazenes, and polysulphate.

Assignment-10

Recommended Books:

1. Lee, J.D. Concise Inorganic Chemistry, ELBS, 1991.

2. Douglas, B.E; Mc Daniel, D.H. & Alexander, J.J. Concepts & Models of Inorganic Chemistry 3rd Ed., John Wiley Sons, N.Y. 1994.

3. Greenwood, N.N. & Earnshaw. Chemistry of the Elements, Butterworth-Heinemann. 1997.

4. Cotton, F.A. & Wilkinson, G. Advanced Inorganic Chemistry, Wiley, VCH, 1999.

5. Rodger, G.E. Inorganic and Solid State Chemistry, Cengage Course India Edition, 2002.

6. Miessler, G. L. & Donald, A. Tarr. Inorganic Chemistry 4th Ed., Pearson, 2010.

7. Atkin, P. Shriver & Atkins' Inorganic Chemistry 5th Ed. Oxford University Press (2010).

8. Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009.

Note: 1 credit theory=10 hrs lecture, 1 credit practice/project=12.5 hrs lab/workshop/field work in a semester

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1474	Oxygen Containing Functional Groups	3-2-1	

Objective

- The aim of this course to introduce basic practical skills to synthesize organic molecules containing functional groups like alcohols, acids, acid derivatives, carbonyl compounds, ethers, etc.
- In addition to that, the course will also help students to understand the reaction mechanism subjects in the later stages of their study.

Course outcome

COs	Course outcomes
CO1	Able to gain broad knowledge in various organic reactions, functional group transformations and their mechanisms
CO2	To identify, design and development of different organic molecular structure for synthesis of required compounds
CO3	Able to enhance the ability to understand the organic chemistry and it will

	help in trouble shooting of organic reaction
CO4	Able to get skill on functional group analysis.
CO5	Able to do research, fabricate the inorganic polymer-based material and interpret the data.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4													3	2	3
CO5													3	2	3

***High-3, Medium-2, Low-1**

Course content

Module-1: Chemistry of Alkyl Halides
Theory (4h)

- Alkyl halides: Methods of preparation of Alkyl halides
- Nucleophilic substitution Unimolecular (SN1) reaction
- Nucleophilic substitution Bimolecular (SN2)reaction and SNi mechanism
- Stereochemical aspects of SN1, SN2 and SNi reactions

Practice 1 (2Hrs)

- Functional group test for Alkyl halides

Assignment (2Hrs)

- Nucleophilic substitution vs. Elimination.
- Effect of solvent on SN1, SN2 reaction mechanism

Module-2: Aryl Halides and Organometallic compounds

Theory (4h)

- Preparation of Aryl halides including preparation from diazonium salts
- Nucleophilic aromatic substitution (SNAr) of aryl halides
- Benzyne mechanism
- Organometallic compounds of Mg and Li and use in the synthesis of organic compounds

Practice 2 (4Hrs)
Preparation of benzene diazonium chloride.
Practice 2 (4Hrs)
Synthesis of chlorobenzene from aniline
Assignment (2Hrs)

- Relative reactivity of alkyl, allyl /benzyl, vinyl and aryl halides towards nucleophilic substitution reactions.
- Synthesis of organic compounds using organometallic compounds of Mg/Li.

Module-3: Alcohols
Theory (4Hrs)

- Preparation of Alcohols
- Properties and relative reactivity of 1°, 2°, 3° alcohols, Bouvaelt-Blanc Reduction
- Preparation and properties of glycols
- Oxidation by periodic acid and lead tetraacetate

Practice 6 (4h)

- Functional group test for alcohol

Assignment (1h)

- Effect of substituents on Pinacol-Pinacolone rearrangement

Module-4: Phenols, ethers and epoxides
Theory (4h)

- Preparation and properties, Acidity and factors affecting the acidic character of Phenol
- Ring substitution reactions, Reimer- Tiemann, Kolbe's-Schmidt Reactions, Fries and Claisen rearrangements with mechanism
- Preparation of Ethers and reactions with acids
- Preparation of Epoxides

Practice 7 (4h)

- Functional group test for Phenol

Assignment (1h)

- Reactions of epoxides with alcohol, ammonia derivatives, and $LiAlH_4$.

Module-5: Carbonyl Compounds-1
Theory (4h)

- Structure, reactivity, and preparation of carbonyl compounds, Nucleophilic additions
- Nucleophilic addition-elimination reactions with ammonia derivatives with the mechanism
- Mechanisms of Aldol and Benzoin condensation, Knoevenagel condensation, Claisen-Schmidt, Perkin, Cannizzaro reaction
- Wittig reaction, Beckmann and Benzil-Benzilic acid rearrangement

Practice 8 (4h)

- Functional group test for Carbonyl compounds

Assignment (1h)

- Draw the mechanism of all the name reactions in this module

Module-6: Carbonyl Compounds-2
Theory (4h)

- Haloform reaction and Baeyer Villiger oxidation, alfa- substitution reactions, oxidations and reductions (Clemmensen, Wolff-Kishner, $LiAlH_4$, $NaBH_4$, MPV, PDC)
- Addition reactions of unsaturated carbonyl compounds: Michael addition, Active methylene compounds: Keto-enol tautomerism.
- Preparation and synthetic applications of diethyl malonate
- Preparation and synthetic applications of ethyl acetoacetate.

Assignment (1h)

- Reactivity order of Carbonyl compounds

Module-7: Carboxylic Acids and their Derivatives and Sulfer containing compounds
Theory (6h)

- Preparation, physical properties, and reactions of monocarboxylic acids
- Typical reactions of dicarboxylic acids, hydroxy acids, and unsaturated acids
- Succinic/phthalic, lactic, malic, tartaric, citric, maleic and fumaric acids
- Preparation and reactions of acid chlorides, anhydrides, esters, and amides
- Preparation and reactions of thiols
- Preparation and reactions thioethers
- Preparation and reactions sulphonic acids

Practice 9 (4h)

- Functional group test for carboxylic acid

Assignment (4h)

- Comparative study of nucleophilic substitution at acyl group -Mechanism of acidic and alkaline hydrolysis of esters
- Draw the mechanism of Claisen condensation, Dieckmann and Reformatsky reactions, Hofmann-bromamide degradation and Curtius rearrangement

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1475	Phase Equilibria and Chemical Kinetics	3-2-1	

Objective

- The course gives idea about the different phases of matter and their equilibria from which the stability and sustainability can be easily predicted.
- Deals with kinetics study of different processes and surface phenomenon like adsorption, chemisorptions etc.

Course outcome After completion of the course, students will be able to

COs	Course outcomes
CO1	Determine the different phases of matter and kinetics of reaction.
CO2	Explain the degrees of freedom of the systems and kinetics of chemical reaction
CO3	Demonstrate the controlling factors that affect the rate of reactions
CO4	Analyze the phase diagram and solve related problems
CO5	Fabricate the inorganic polymer-based material

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	3	1	1	2	3	2	3	2	1	2	1	2	1

CO2	2	3	3	2	2	2	3	1	2	1	2	2	2	3	2
CO3	3	2	2	3	3	1	1	2	3	1	1	2	3	2	3
CO4	3	3	3	2	2	3	2	1	2	3	2	1	2	1	2
CO5	2	3	1	2	2	2	1	2	1	2	3	3	3	2	3

***High-3, Medium-2, Low-1**

Course content

Module 1: **Phase Equilibria-1 (8 hrs)**

Definitions of phase, Concept of phases and Components; Degrees of freedom; Derivation of Gibbs Phase Rule for Non-reactive and reactive systems; Clausius-Clapeyron equation; Applications to solid-liquid, Liquid-vapour and Solid- vapour equilibria.

Module II: **Phase Equilibria-2 (9 hrs)**

Definition of phase diagram; Phase diagram for one component systems – water Phase diagram for one component systems – CO_2, with applications; Liquid-vapour equilibrium for two component systems; Phase diagrams for systems of solid-liquid equilibria involving eutectic, congruent and incongruent melting points, solid solutions, Three component systems, water-chloroform-acetic acid system, triangular plots.

Module III: **Phase Equilibria-3 (8 hrs)**

Gibbs- Duhem-Margules equation and its derivation; Applications to fractional distillation of binary miscible liquids (ideal and nonideal); Azeotropes, lever rule; Partial miscibility of liquids, CST; Miscible pairs, steam distillation; Nernst distribution law: its derivation and applications.

Practice 1

Study the distribution law of Acetic acid and benzoic acid by water and carbon tetra chloride

Module IV: **Chemical Kinetics-1 (12 hrs)**

Order and molecularity of a reaction; Rate laws in terms of the advancement of a reaction, rate constants; Differential and integrated form of rate expressions up to second order reactions; Pseudo first order reactions; Experimental methods of the determination of rate laws; Determination of order of a reaction by half-life and differential method; Kinetics of complex

reactions (integrated rate Expressions up to first order only).

Practice 2
Kinetics Study on the Reaction between Sodium Thiosulphate and Hydrochloric Acid

Practice 3
Acid hydrolysis of methyl acetate with hydrochloric acid

Module V: **Chemical Kinetics-2 (9 hrs)**

Opposing reactions and parallel reactions; Consecutive reactions and their differential rate equations (steady-state approximation in reaction mechanisms); Chain reactions; Temperature dependence of reaction rates; Arrhenius equation; activation energy; Unimolecular Reactions; Collision theory of reaction rates, Lindemann mechanism, qualitative treatment of the theory of absolute reaction rates.

Practice 4
Effect of ionic strength on rate of persulphate iodide reaction

Module VI: **Catalysis (6 hrs)**

Types of catalyst, specificity and selectivity, mechanisms of catalyzed reactions at solid surfaces; Salt effects; effect of particle size and efficiency of nanoparticles as catalysts; Enzyme catalysis, Michaelis-Menten mechanism; Acid-base catalysis, turn-over number.

Module VII: **Surface Chemistry (12 hrs)**

Physical adsorption, chemisorptions; Freundlich and Langmuir adsorption isotherms; Multilayer adsorption and BET isotherm; Gibbs adsorption isotherm and surface excess, nature of adsorbed state.

Practice 5
Verify the Freundlich for adsorption of acetic acid on activated charcoal

Practice 6
Verify the Langmuir isotherm for adsorption of acetic acid on activated charcoal

Text Books:

COs	Course outcomes

1. Peter Atkins & Julio De Paula, Physical Chemistry 10th Ed., Oxford University Press (2014).
2. Castellan, G. W. Physical Chemistry, 4th Ed., Narosa (2004).
3. Levine, I. N. Physical Chemistry 6th Ed., Tata McGraw-Hill (2011).

Reference Books:

1. Engel, T. & Reid, P. *Physical Chemistry 3rd Ed.,* Prentice-Hall (2012).
2. Zumdhal, S.S. *Chemistry concepts and applications* Cengage India (2011).
3. Ball, D. W. *Physical Chemistry* Cengage India (2012).
4. Mortimer, R. G. *Physical Chemistry 3rd Ed.,* Elsevier: NOIDA, UP (2009).
5. Metz, C. R *Physical Chemistry 2nd Ed.,* Tata McGraw-Hill (2009).

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1476	Coordination Chemistry	3-2-1	

Objective

- To Know the basic of coordination chemistry, bio-inorganic chemistry and aimed at advanced knowledge in the field of industrial chemistry
- To be able to describe the stability of metal complexes by the use of formation constants and to calculate thermodynamic parameters from them
- To able to know the bonding and structure of coordination compounds and their applications

Course outcome

After completion of the course students will be able to

t

CO1	Demonstrate advanced knowledge and understanding in coordination compound
CO2	Understand the concepts of metal ligand bonding in transition complex compounds
CO3	Describe role of different metal ions in biological system
CO4	Recognize the biological reaction alkali and alkaline earth metals, nitrogen fixation, hemoglobin and myoglobin
CO5	Explain the role of Inorganic ―substances in living systems and the use of metal ions in medicinal therapy and diagnosis field

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4													3	2	3
CO5													3	2	3

***High-3, Medium-2, Low-1**

Course Content

Module-I (7Hours)

Coordination Chemistry

Werner's theory, Valence bond theory (inner and outer orbital complexes), Electro neutrality principle and back bonding.

Practice

- To Synthesis and Characterization of Metal Complexes with Schiff Base Ligands
- Estimation of nickel (II) using Dimethylglyoxime (DMG).

Module-II (11Hours)

Crystal Field Theory

Crystal field theory, measurement of 10 Dq (Δo), CFSE in weak and strong fields, pairing energies, factors affecting the magnitude of 10 Dq (Δo, Δt). Octahedral vs. tetrahedral coordination, tetragonal distortions from octahedral geometry Jahn-Teller theorem, square planar geometry. Qualitative aspect of Ligand field and MO Theory.

Assignment -1

Module-III(8 Hours)

Nomenclature of Coordination Compounds

IUPAC nomenclature of coordination compounds, isomerism in coordination compounds. Stereochemistry of complexes with 4 and 6 coordination numbers. Chelate effect of polynuclear complexes, Labile and inert complexes.

Assignment-2

Practice 1

- To Prepare of Ferrous ammonium salt and estimation of Nitrogen
- **Practice 2**
- Preparation and Analysis of Potassium Trioxalatoferrate (III) Trihydrate

Module-IV(11 Hours)

Transition Elements:

General group trends with special reference to electronic configuration, colour, variable valency, magnetic and catalytic properties, ability to form complexes. Stability of various oxidation states and e.m.f. (Latimer & Bsworth diagrams). Difference between the first, second and third transition series. Chemistry of Ti, V, Cr Mn, Fe and Co in various oxidation states (excluding their metallurgy)

Assignment-3

Assignment-4

Practice 3

To estimate the amount of barium in the whole of the given solution of barium chloride

Practice 4
Estimation of Iron in Hematite ore solution

Module-V (9hours)

Lanthanoids and Actinoids:

Electronic configuration, oxidation states, colour, Spectral and magnetic properties, lanthanide contraction, separation of lanthanides (ion-exchange method only),Actinides elements and properties

Assignment-5

Module-VI(11Hours)

Thermodyanamics and Kinetics

Introduction, Stability and lability of complexes, Steps Involved in Formation of a Complex MLn, Stepwise and Overall Stability Constants, Explanation of lability and inertness of octahedral complexes, Factors Affecting the Stability and Lability of Complexes, Methods for detection of complex Formation, Experimental determination of stability constant and composition of a complex,

Assignment-6

Assignment-7

Assignment-8

Practice 5
Determination of Iron from a supplied waste water sample by redox and spectroscopic method

Module-VII(7Hours)

Reactions in Square Planar Complexes

Introduction, The trans effect and its applications, Theories for explaining trans effect, Mechanism of substitution reactions, Factors affecting the rate of substitution reactions in square planar complexes,

Assignment-9

Assignment-10

Recommended Books:

1. Purcell, K.F & Kotz, J.C. Inorganic ChemistryB. Saunders Co, 1977.
2. Huheey, J.E., Inorganic Chemistry, Prentice Hall, 1993.
3. Lippard, S.J. & Berg, J.M. Principles of Bioinorganic Chemistry Panima Publishing Company 1994.
4. Cotton, F.A. & Wilkinson, G, Advanced Inorganic Chemistry Wiley-VCH, 1999
5. Bassolo, F, and Pearson, R.C. Mechanisms of Inorganic Chemistry, John Wiley & Sons, NY, 1967.
6. Greenwood, N.N. & Earnshaw A. Chemistry of the Elements, Butterworth-Heinemann, 1997.

Recommended Books For Practice:

1.Mendham, J., A. I. Vogel's Quantitative Chemical Analysis 6th Ed., Pearson, 2009

Note: 1 credit theory=10 hrs lecture, 1 credit practice/project=12.5 hrs lab/workshop/field work in a semester

Course Outline

Code	Course Title	T-P-P	Prerequisite
CUTM1477	Heterocyclic Chemistry	3-2-1	

Course Objectives

- To introduce students to Nitrogen containing functional groups and their application in organic conversions and related mechanisms.
- Students are also expected to learn about structure, synthesis, reactivity of important heterocyclic compounds and polycyclic aromatic hydrocarbons.
- To familiarize students about different classes of N-based naturally occurring important alkaloid and terpenoid compounds, their structures, synthesis and reactivity.

Course Outcomes

After completion of this course Students will be able to

COs	Course outcomes

CO1	Remember the properties and synthesis of heterocyclic compounds and basic knowledge in bio-organic chemistry
CO2	Understand the concepts and classification of different heterocyclic compounds
CO3	Apply the organic chemistry principle for trouble shooting of organic reaction
CO4	Analyze the extraction, purification and separation techniques for synthesis various organic molecules.
CO5	Fabricate different heterocyclic structures and can create new molecules in the laboratory.

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4		2	2	1	1								3	2	3
CO5		1	2	2	2								3	2	3

***High-3, Medium-2, Low-1**

Course Content

Module-I (5 Hrs.)

Nitrogen Containing Functional Groups: Nitro compounds, nitriles and isonitriles

1.1 Structure and Preparation of nitroalkanes and nitroarenes
1.2 Properties and reactions of nitroalkanes

1.3 Properties and reactions of nitroarenes
1.4 Structure, Preparation and properties of nitriles and isonitriles

Assignment 1/Flip class: Important Nitro aromatic compounds and their uses

Module-II (8 Hrs.)

Nitrogen Containing Functional Groups:Amines
2.1 Preparation of primary amines: Reduction of nitro compounds, Hofmann ammonolysis, Hofmann degradation, Gabriel phthalimide synthesis.
2.2 Preparation of secondary and tertiary amines: Aminolysis of alkyl halides, Reductive amination of aldehydes and ketones, Ullmann reaction
2.3 Properties of amines, Basicity, Effect of substituent and solvent on basicity.
2.4 Important reactions of amines: Alkylation, acylation, Carbylamine reaction
2.5 Important reactions of amines: Mannich reaction, Hoffmann's exhaustive methylation, Hofmann-elimination reaction
2.6 Diazonium Salts: Structure, Preparation and reactions/applications

Assignment 2/Flip class:Electrophilic substitution reactions of aryl amines

Assignment 3/Flip class:Distinction between 1°, 2° and 3°amines with Hinsberg reagent and nitrous acid.

Laboratory Practices (Module I and II): (13.30 hrs.)

Practice 1. **Detection of Nitrogen**
Practice 2. **Detection of Sulphur**
Practice 3. **Detection of Halogens**
Practice 4. **Functional group test for: Nitro groups**
Practice 5. **Functional group test for: Amine groups**
Practice 6. **Functional group test for: Amide groups**

Module-III (8 Hrs.)
Five membered Heterocyclic Compounds containing one heteroatom:
3.1 Classification, nomenclature and structure of pyrrole, furan and thiophene (5-numbered) and pyridine (6-membered)
3.2 Molecular orbital pictures and aromaticity in of pyrrole, furan and thiophene and pyridine
3.3 Synthesis, reactions and mechanism of substitution reactions of: Furan
3.4 Synthesis of Pyrrole: Knorr pyrrole synthesis, Paal-Knorr synthesis, Hantzsch synthesis.
3.5 Reactions and mechanism of substitution reactions of Pyrrole
3.6 Derivatives of furan: Furfural and furoic acid.

Assignment 4/Flip class:Synthesis and Properties of thiophene

Assignment 5/Flip class:Acidic and basic character of Pyrrole, Furan and Thiophene

Module-IV (6 Hrs.)
Six membered and condensed Heterocyclic Compounds:
4.1 Structure, synthesis and properties of Pyridine (Hantzsch synthesis), Pyrimidine
4.2 Structure elucidation of indole, Fischer indole synthesis and Madelung synthesis)
4.3 Structure elucidation of quinoline and isoquinoline
4.4 Skraup synthesis, Friedlander's synthesis

Assignment 6/Flip class: Knorr quinoline synthesis, Doebner- Miller synthesis

Assignment 7/Flip class:Bischler-Napieralski reaction, Pictet-Spengler reaction, Pomeranz-Fritsch Reaction.

Module-V (5 Hrs.)
Polynuclear Hydrocarbons:
5.1 Preparation and structure elucidation of naphthalene
5.2 Reactions of naphthalene
5.3 Important derivatives of naphthalene
5.4 Preparation, structure elucidation and important derivatives of anthracene.

Assignment 8/Flip class:Preparation, structure elucidation and properties of phenanthrene

Module-VI (6 Hrs.)
Alkaloids
6.1 Natural occurrence, Isolation and their physiological action
6.2 General structural features, experimental determination
6.3 Hoffmann's exhaustive methylation, Emde's modification.
6.4 Structure elucidation and synthesis of Hygrine
6.5 Structure elucidation and synthesis of Nicotine.

Assignment 9/Flip class;Medicinal importance of Nicotine, Hygrine, Quinine, Morphine, Cocaine, and Reserpine

Module-VII (3 Hrs.)
Terpenoids:
7.1 Occurrence, classification, isoprene rule;
7.2 Elucidation of structure and synthesis of Citral, Neral

Assignment 10/Flip class:Elucidation of structure and synthesis of α- terpineol.

Text Books :

1. A Textbook of Organic Chemistry – III, M. K. Jain, S. C. Sharma, Amita, Vishal Publishing Co.
2. Kalsi, P. S. Textbook of Organic Chemistry 1st Ed., New Age International (P) Ltd. Pub.
3. Morrison, R. T. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education)
4. Clayden, J.; Greeves, N.; Warren, S.; Wothers, P.; Organic Chemistry, Oxford University Press.

Reference Books :

1. Acheson, R.M. Introduction to the Chemistry of Heterocyclic compounds, John Welly & Sons (1976).
2. Heterocyclic Chemistry, Fifth Edition, J. A. Joule, K. Mills and G. F. Smith
3. Singh, J.; Ali, S.M. & Singh, J. Natural Product Chemistry, PragatiParakashan (2010).

Course Outline

Code	Course Title	T-P-P (Credit)	Prerequisite
CUTM1478	Electrochemistry	3-2-1	

Objective

- To Know the basic of ions, electrolyte, movement of ions, electrochemistry
- To know how the ionic movements are related to different other fields such as thermodynamics.
- Also, this course will help students to garner basic knowledge on novel energy storage devices

Course outcome

After completion of the course students will be able to

COs	Course outcomes
CO1	Define the ionic, and electrical properties of sample
CO2	Explain the behavior of electrolytes in solutions
CO3	Able to gain skill on the handling of conductometer and to Interpret the results of electrolytic conductance in different solution

CO4	Differentiate the conductometric behavior of different electrolytic solution based on titration
CO5	Develop different electrochemical device for energy storage

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4		2	1	1	1								3	2	3
CO5		1	2	1	1								3	2	3

***High-3, Medium-2, Low-1**

Course content

Module-I: Conductance I(12 hrs)

Arrhenius theory of electrolytic dissociation. Conductivity, equivalent and molar conductivity and their variation with dilution for weak and strong electrolytes. Molar conductivity at infinite dilution. Kohlrausch law of independent migration of ions.

Assignment 1: Molar conductivity and their variation with dilution for weak and strong electrolytes.

Assignment 2: Ostwald's dilution law.

Module-II: Conductance II (5hrs)

Walden’s rules, Debye-Huckel-Onsager equation. Ionic mobility and their determinations, transference numbers and their relation to ionic mobilities, determination of transference numbers using Hittorf and Moving Boundary methods.

Assignment 3: Walden's rules, Debye-Huckel-Onsager equation

Assignment 4: Ionic mobility and their determinations, transference numbers and their relation to ionic mobilities.

Module-III: Conductance III (6 hrs)

Applications of conductance measurement: (i) degree of dissociation of weak electrolytes, (ii) ionic product of water (iii) solubility and solubility product of sparingly soluble salts, (iv) conductometric titrations, and (v) hydrolysis constants of salts.

Practice 1: Conductometric titrations of strong acid and strong base

Practice 2: Conductometric titrations of strong acid and weak base

Practice 3: Conductometric titrations of weak acid and strong base

Practice 4: Conductometric titrations of weak acid and weak base

Practice 5: Conductometric titrations of mixture of strong and weak acid with a strong base

Assignment 5: Applications of conductance measurement for determining hydrolysis constants of salts.

Module-IV: Electrochemistry I (17 hrs)

Quantitative aspects of Faraday's laws of electrolysis, applications of electrolysis in metallurgy and industry. Electrochemical series, rules of oxidation/reduction of ions based on half-cell potentials. Nernst equation; Standard electrode (reduction) potential and its application to different kinds of half-cells. Electromotive force of a cell and its measurement. Chemical cells, reversible and irreversible cells with examples.

Assignment 6: Quantitative aspects of Faraday's laws of electrolysis

Assignment 7: Applications of electrolysis in metallurgy and industry.

Module-V: Electrochemistry II(8 hrs)

Application of EMF measurements in determining (i) free energy, enthalpy and entropy of a cell reaction, (ii) equilibrium constants, and (iii) pH values, using hydrogen, quinone-hydroquinone, and glass electrodes. Concentration cells with and without transference, liquid junction potential;

determination of activity coefficients and transference numbers. Qualitative discussion of potentiometric titrations (acid-base, redox, precipitation).

Assignment 8: Concentration cells without transference.

Module-VI: Electrical & Magnetic Properties of Atoms and Molecules(6 Hrs)

Basic ideas of electrostatics, dielectric constant, Dipole moment and molecular polarizabilities and their measurements. Basics of diamagnetism, paramagnetism. Magnetic susceptibility and its measurement.

Assignment 9: Basic ideas of electrostatics, dielectric constant.

Module-VII: Energy Storage and Conversion(12 Hrs.)

Fundamentals on Li ion batteries, basic principle and types of fuel cells, theory of solar cells and types of solar cells, basic principles on super capacitors and types of super capacitors.

Assignment 10: Needs for energy storage and storage alternatives.

Text Books:

1. Atkins, P.W & Paula, J.D. Physical Chemistry, 10th Ed., Oxford University Press (2014).
2. Castellan, G. W. Physical Chemistry 4th Ed., Narosa (2004).
3. Mortimer, R. G. Physical Chemistry 3rd Ed., Elsevier: NOIDA, UP (2009).
4. Barrow, G. M., Physical Chemistry 5th Ed., Tata McGraw Hill: New Delhi (2006).
5. Engel, T. & Reid, P. Physical Chemistry 3rd Ed., Prentice-Hall (2012).
6. Rogers, D. W. Concise Physical Chemistry Wiley (2010).

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1479	Biomolecules	3-2-1	

Objective

- To Know the basic of Bio-Organic chemistry and its application in industry
- This course gives idea about the structure of different bases of nucleic acid, DNA and RNA.
- Number of amino acids, there functions and the peptide bond that connect di, tri and

polypeptides.
- To study about proteins, lipid and carbohydrates

Course outcome

After completion of this course students will be able to

COs	Course outcomes
CO1	Understand the essential and nonessential amino acids and can predict how their ionic charges change with pH.
CO2	Remember the classification of different biomolecules on the basis of their functionality in living organisms
CO3	One can apply this fundamental chemistry in making new drug molecules against various diseases.
CO4	This course will help students to analyze different Pharmaceutical Compounds for various use such as antipyretics, analgesics, anti malaria or antibiotics applications.
CO5	They can evaluate Bio-imaging and Bio-Medical science and can create diagnostic kit and various biomedical device fabrication .

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	-	-	-						2		3	2	3
CO2	3	2	2	1	-						2		3	2	3
CO3	-	3	2	1	2						2		3	2	3
CO4	-	3	2	1	2						2		3	2	3
CO5	-	3	2	1	2						2		3	2	3

***High-3, Medium-2, Low-1**

Course content

Module I: Nucleic Acids (8 h)
Nucleic Acids: Introduction of Nucleic Acids, Components of nucleic acids, Nucleosides and nucleotides, Structure, synthesis and reactions of: Adenine, Guanine, Cytosine, Uracil and Thymine, Structure of polynucleotides.

Practice 1: Isolation and characterization of DNA from cauliflower

Practice 2: Isolation and characterization of DNA from onion

Assignment 1: Structure of polynucleotides

Module II: Amino acids, peptide and proteins (13 h)
Amino acids, peptide and proteins: Introduction to amino acid, peptide and protein, Classification of amino acid and peptides, α-Amino Acids: Synthesis, ionic properties and reactions, Zwitterions, pKa values, isoelectric point and electrophoresis, Study of peptides: determination of their primary structures-end group analysis, methods of peptide synthesis (Synthesis of peptides using N-protecting, C-protecting and C-activating groups -Solid-phase synthesis)

Practice 3: Estimation of glycine by Sorenson's formalin method

Practice 4: Estimation of proteins by Lowry's method.

Assignment 2: Reactions, Zwitterions, pKa values, isoelectric point and electrophoresis

Assignment 3: α-Amino Acids: Study of peptides: determination of their primary structures-end group analysis

Module III:Enzymes (6 h)
Enzymes: Introduction, classification and characteristics of enzymes, Salient features of active site of enzymes, Mechanism of enzyme action, Factors affecting enzyme action, Coenzymes and Cofactors and their role in biological reactions, specificity of enzyme action (including stereospecificity), Enzyme inhibitors: Importance of enzyme inhibition phenomenon of inhibition (competitive, uncompetitive and non-competitive inhibition including allosteric inhibition).

Assignment 4: Enzyme inhibitors: Importance of enzyme inhibition phenomenon of inhibition (non-competitive inhibition including allosteric inhibition)

Module IV: Lipids (10 h)
Lipids: Introduction to oils and fats (Properties and functions), Classes of Lipids (common fatty acids present in oils and fats, Examples of diff. Lipids), Hydrogenation of fats and oils, Saponification value, acid value, iodine number, Reversion and rancidity

Practice 5: Saponification value of an oil or a fat

practice 6: Determination of Iodine number of an oil/ fat.

Module V: Concept of Energy in Biosystems (8 h)
Concept of Energy in Biosystems: Introduction to metabolism (catabolism, anabolism), ATP: ATP hydrolysis and free energy change, Biological redox systems: NAD+, FAD , Conversion of food to energy, Outline of catabolic pathways of carbohydrate- glycolysis, Fermentation and Krebs cycle, Catabolic pathways of fat and protein, Metabolic pathways of protein, fat and carbohydrate

Practice 7: Extraction of starch from potatoes

Assignment 5: Outline of catabolic pathways of carbohydrate-glycolysis , Fermentation and Krebs cycle

Assignment 6: Catabolic pathways of fat and protein, Metabolic pathways of protein, fat and carbohydrate

Module VI: Pharmaceutical Compounds (9 h)
Pharmaceutical Compounds: Structure and Importance: Classification, structure and therapeutic uses of antipyretics: Paracetamol (with synthesis) , Analgesics: Ibuprofen (with synthesis), Antimalarials: Chloroquine, Antibiotics and detailed study of chloramphenicol, Medicinal values of curcumin (haldi), azadirachtin (neem), vitamin C and antacid (ranitidine)

Assignment 7: Medicinal values of curcumin (haldi), azadirachtin (neem), vitamin C and antacid (ranitidine)

Module VII (12 h)
Bio-imaging and Bio-Medical science: Introduction to basic cell biology, Optical fluorescent probe and photophysical property, Optical probe for bio-imaging application, Protein labelling, Diagnostic kit and Device fabrication

Assignment 8: Diagnostic kit development

Assignment 9: Device fabrication

Reference Books:

1. Berg, J.M., Tymoczko, J. L. and Stryer, L. (2006) Biochemistry. VIth Edition. W.H. Freeman and Co.
2. Nelson, D. L., Cox, M. M. and Lehninger, A. L. (2009) Principles of Biochemistry. IV Edition. W.H. Freeman and Co.
3. Murray, R. K., Granner, D. K., Mayes, P. A. and Rodwell, V.W. (2009) Harper's Illustrated Biochemistry. XXVIII edition. Lange Medical Books/ McGraw-Hill.

Note: 1 credit theory=10 hrs lecture, 1 credit practice/project=12.5 hrs lab/workshop/field work in a semester

Course Outline

Nomenclature

Code	Course Title	T-P-P (Credit)	Prerequisite
CUTM1480	Quantum Chemistry and Spectroscopy	3-2-1	

Objective

- To impart knowledge about quantum mechanical principles and understanding as well as predicting different microscopic phenomena.
- To understand the covalent nature of bonding and their theoretical background and correlation to practical aspects.
- To have a sound understanding of different spectroscopic techniques and photochemistry.

Course outcome

After completion of the course students will be able to

COs	Course outcomes
CO1	Students will gain knowledge on different aspects of quantum chemistry and its applicability in different bonding behavior of atoms.
CO2	Differentiate between classical and quantum mechanics to understand behavior of microscopic objects.

CO3	Develop skill to identify, predict and analyze different types of interactions of radiation with matter which is crucial for research purposes.
CO4	Gain problem-solving skills from spectroscopic data
CO5	Such knowledge on quantum mechanics and spectroscopy will help them to do research and predict/interpret data.

Course Outcome to Program Outcome Mapping:

COs	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
CO1	3	2	2	3	3								3	2	3
CO2	3	3	3	3	3								3	2	3
CO3	-	3	2	3	3								3	2	3
CO4													3	2	3
CO5													3	2	3

***High-3, Medium-2, Low-1**

Course content

Module I Quantum Chemistry (18 hours)

Postulates of quantum mechanics, quantum mechanical operators, Schrödinger equation and its application to free particle and "particle-in-a-box" (rigorous treatment), quantization of energy levels, Heisenberg Uncertainty principle; wavefunctions, probability distribution functions, nodal properties, Extension to two and three dimensional boxes, separation of variables, degeneracy. Qualitative treatment of simple harmonic oscillator model of vibrational motion: Setting up of Schrödinger equation and discussion of solution and wavefunctions. Vibrational energy of diatomic molecules and zero-point energy.

Assignment 1: Apply Schrodinger's equation to hydrogen atom

Angular momentum: Commutation rules, quantization of square of total angular momentum and z-component. Rigid rotator model of rotation of diatomic molecule. Schrödinger equation, transformation to spherical polar coordinates. Separation of variables. Spherical harmonics. Discussion of solution. Qualitative treatment of hydrogen

atom and hydrogen-like ions: setting up of Schrödinger equation in spherical polar coordinates, radial part, quantization of energy (only final energy expression). Average and most probable distances of electron from nucleus.Setting up of Schrödinger equation for many-electron atoms (He, Li). Need for approximation methods. Statement of variation theorem and application to simple systems(particle-in-a-box, harmonic oscillator, hydrogen atom).

Assignment 2 : Draw molecular orbital energy diagram for LiH molecule

Module II

Chemical bonding (5 hours)

Covalent bonding, valence bond and molecular orbital approaches, LCAO-MO treatment of H_2^+. Bonding and antibonding orbitals. Qualitative extension to H_2. Comparison of LCAO-MO and VB treatments of H_2 (only wavefunctions, detailed solution not required) and their limitations.

Refinements of the two approaches (Configuration Interaction for MO, ionic terms in VB). Qualitative description of LCAO-MO treatment of
homonuclear and heteronuclear diatomic molecules (HF, LiH). Localised and non-localised molecular orbitals treatment of triatomic (BeH_2, H_2O) molecules. Qualitative MO theory and its application to AH_2 type molecules.

Assignment : Give the detailed LCAO treatment to Hydrogen atom

Module III (3 hours)

Introduction to Molecular Spectroscopy:

Interaction of electromagnetic radiation with molecules and various types of spectra; Born-Oppenheimer approximation.

Rotation spectroscopy: Selection rules, intensities of spectral lines, determination of bond lengths of diatomic and linear triatomic molecules, isotopic substitution.

Assignment 4: Find the condition of maximum intensity for rotational spectra.

Module IV

Vibrational spectroscopy (7 hours)

Classical equation of vibration, computation of force constant,amplitude of diatomic molecular vibrations, anharmonicity, Morse potential, dissociation energies, fundamental frequencies, overtones, hot bands, degrees of freedom for polyatomic molecules, modes of vibration, concept of group frequencies. Vibration-rotation spectroscopy: diatomic vibrating rotator, P, Q, R branches.

Assignment 5: Draw and explain Moorse potential energy diagram

Module V

Raman and Electronic Spectroscopy (10 hours)

Raman spectroscopy: Qualitative treatment of Rotational Raman effect; Effect of nuclear spin, Vibrational Raman spectra, Stokes and anti-Stokes lines; their intensity difference, rule of mutual exclusion.

Electronic spectroscopy: Franck-Condon principle, electronic transitions, singlet and triplet states, fluorescence and phosphorescence, dissociation and predissociation, calculation of electronic transitions of polyenes using free electron model.

Assignment : Give mathematical deduction of Raman effect.

Practice 1

Study the 200-500 nm absorbance spectra of $KMnO_4$ and $K_2Cr_2O_7$ (in 0.1 M H_2SO_4) and determine the λ_{max} values. Calculate the energies of the two transitions in different units (J mol^{-1}, kJ mol^{-1}, cm^{-1}, eV).

Practice 2

Study the pH-dependence of the UV-Vis spectrum (200-500 nm) of $K_2Cr_2O_7$.

Practice 3

Record the 200-350 nm UV spectra of the given compounds (acetone, acetaldehyde,2-propanol, acetic acid) in water. Comment on the effect of structure on the UVspectra of organic compounds.

Module VI (6 hours)

Nuclear Magnetic Resonance (NMR) spectroscopy: Principles of NMR spectroscopy, Larmor precession, chemical shift and low resolution spectra, different scales, spin-spin coupling and high resolution spectra, interpretation of PMR spectra of organic molecules.

Electron Spin Resonance (ESR) spectroscopy: Its principle, hyperfine structure, ESR of simple radicals.

Assignment 6: Deduce the number of ESR peaks for methyl radical.

Module VII Photochemistry (16 hours)

Characteristics of electromagnetic radiation, Lambert-Beer's law and its limitations, physicalsignificance of absorption coefficients.

Laws, of photochemistry, quantum yield, actinometry, examples of low and high quantum yields, photochemical equilibrium and the differentialrate of photochemical reactions, photosensitised reactions, quenching. Role of photochemical reactions in biochemical processes, photostationary states, chemiluminescence.

Practice 4

Verify Lambert-Beer's law and determine the concentration of $CuSO_4$/$KMnO_4$/$K_2Cr_2O_7$ in a solution of unknown concentration

Practice 6
Determine the amount of iron present in a sample using 1,10-phenathroline.

Practice 7
Determine the dissociation constant of an indicator (phenolphthalein).

Practice 8
Study the kinetics of interaction of crystal violet/ phenolphthalein with sodium hydroxide.

Reference Books:

• Banwell, C. N. & McCash, E. M. Fundamentals of Molecular Spectroscopy 4th Ed. Tata McGraw-Hill: New Delhi (2006).
• Chandra, A. K. Introductory Quantum Chemistry Tata McGraw-Hill (2001).
• House, J. E. Fundamentals of Quantum Chemistry 2nd Ed. Elsevier: USA (2004).
• Lowe, J. P. & Peterson, K. Quantum Chemistry, Academic Press (2005).
• Kakkar, R. Atomic & Molecular Spectroscopy, Cambridge University Press (2015).

CHEMISTRY PRACTICAL-C XII LAB

Practice 1. Verify Lambert-Beer's law and determine the concentration of $CuSO_4$/$KMnO_4$/$K_2Cr_2O_7$ in a solution of unknown concentration

Practice 2. Determine the concentrations of $KMnO_4$ and $K_2Cr_2O_7$ in a mixture.

Practice 3. Study the kinetics of iodination of propanone in acidic medium.

Practice 4. Determine the amount of iron present in a sample using 1,10-phenathroline.

Practice 5. Determine the dissociation constant of an indicator (phenolphthalein).

Practice 6. Study the kinetics of interaction of crystal violet/ phenolphthalein with sodium hydroxide.

Reference Books

• Khosla, B. D.; Garg, V. C. & Gulati, A., *Senior Practical Physical Chemistry,* R. Chand & Co.: New Delhi (2011).

• Garland, C. W.; Nibler, J. W. & Shoemaker, D. P. *Experiments in Physical Chemistry 8th Ed.;* McGraw-Hill: New York (2003).
• Halpern, A. M. & McBane, G. C. *Experimental Physical Chemistry 3rd Ed.;* W.H. Freeman & Co.: New York (2003).

Note: 1 credit theory=10 hrs lecture, 1 credit practice/project=12.5 hrs lab/workshop/field work in a semester

Course Outline

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1481	**Organometallic Chemistry**	**3-2-1**	

Objective

- To learn about the 18e rule and its violation.
- To identify the basic concept, terms, and important events in the development of organometallic chemistry.
- To learn methods, including spectroscopy techniques, used to determine the structure of organometallic complexes and to probe reaction mechanism.
- To develop an appreciation for the scope, diversity, and application of organometallic chemistry.
- To learn about the common organometallic reactions and to be able to draw reasonable reaction mechanisms.

Course outcome

After completion of the course students will be able to

COs	Course outcomes
CO1	Develop understanding of the fundamental principles of organotransition-metal chemistry and know how chemical properties are affected by metals and ligands.
CO2	Gain knowledge about structure and bonding issues to understand the stability and reactivity of simple organometallic complexes.
CO3	Have insight into the use of modern methods to characterize organometallic compounds.
CO4	Understand fundamental reaction types and mechanisms and how to combine these to understand efficient catalytic processes
CO5	Know important applications of organometallic homogeneous catalysis in the production of large-scale (bulk) and smaller-scale (fine chemicals) production.

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	2	3	3	1	3	2				3	3	2	3
CO2	3	3	3	2	3	1	2	1				3	3	2	3
CO3	-	3	2	3	3	1	1	1				2	3	2	3
CO4	2	2	3	3	3	2	2	2				3	3	2	3
CO5	2	2	2	3	3	2	3	2					3	2	3

***High-3, Medium-2, Low-1**

Course content

Module-I (8hours)

Organometallic Compounds Definition and classification of organometallic compounds,The

different organometallic Compounds the basis of bond type.Concept of hapticity of organic ligands.

Assignment -1/Flip-class-1 : Concept of hapticity of organic ligands

Practice -1 : Estimation of Fe(II) with $K_2Cr_2O_7$

Module-II (12hours)

Metal carbonyls: 18 electron rule electron count of mononuclear, polynuclear and substituted metal carbonyls of 3d series.General methods of preparation (direct combination, reductive carbonylation, thermal and photochemical decomposition) of mono and binuclear carbonyls of 3d series.Organ metallic Compounds with metal-metal bond Cluster.Structures of mononuclear and binuclear carbonyls of Cr, Mn, Fe, Co and Ni using VBT.Pi-acceptor behaviour of CO (MO diagram of CO to be discussed.

Assignment -2/Flip-class-2 : Pi-acceptor behaviour of CO (MO diagram of CO to be discussed)

Practice -2 : To detect the presence of a cation and anion in a given inorganic mixture.

Module-III (9hours)

Synergic effect organometallic compounds having ligands with back bonding as metal carbonyl. metal carbenes, metal nitrosyls, IR data to explain extent of back bonding.

Assignment -3/Flip-class-3 : IR data to explain extent of back bonding

Practice -4 : Qualitative analysis of cations part in a supplied salt.

Module-IV(11hours)

Zeise's salt: Preparation and structure, evidences of synergic effect and comparison of synergic effect with that in carbonyls.Metal Alkyls: Important structural features of methyl lithium (tetramer) and trialkylaluminium (dimer), concept of multicentre bonding in these compounds. Role of triethylaluminium in polymerisation of ethene (Ziegler - Natta Catalyst). Ferrocene: Preparation and reactions (acetylation, alkylation, metallation, Mannich Condensation). Structure and aromaticity.Comparison of aromaticity and reactivity with that of benzene.

Assignment -4/Flip-class-4 : Concept of multicentre bonding in these compounds.

Assignment -5/Flip-class-5: Structure and aromaticity.Comparison of aromaticity and reactivity with that of benzene.

Assignment -6/Flip-class-6 : Preparation and reactions (acetylation, alkylation, metallation, Mannich Condensation).

Practice -5 : Silver nanoparticles synthesis from plant extract

ModuleV(5hours)

Reaction Kinetics and Mechanism Introduction to inorganic reaction mechanisms.Substitution reactions in square planar complexes, Transeffect, theories of trans effect,Mechanism of nucleophilic substitution in square planar complexes.

Module VI (6 hours)

Thermodynamic and Kinetic stability, Kinetics of octahedral substitution, Ligand field effects and reaction rates, Mechanism of substitution in octahedral complexes. Catalysis by Organometallic Compounds

Assignment -7/Flip-class-7 : Mechanism of substitution in octahedral complexes. Catalysis by Organometallic Compounds

Assignment -8/Flip-class-8 : Kinetics of octahedral substitution ,Ligand field effects and reaction rates

Practice 6: Estimation of free alkali present in different soaps/detergents

Module-VII(8hours)

Study of the following industrial processes and their mechanism: 1. Alkene hydrogenation (Wilkinsons Catalyst),Hydroformylation (Co salts),Wacker Process, Synthetic gasoline (Fischer Tropsch reaction).

Assignment -9/Flip-class-9: Study of the following industrial processes and their mechanism: 1. Alkene hydrogenation (Wilkinsons Catalyst),Hydroformylation (Co salts)

Assignment -10/Flip-class-10: Synthetic gasoline (Fischer Tropsch reaction) ,Wacker Process.

Practice -7: Estimation of Fe(II) and oxalic acid using standardized KMnO4 solution.

Text Books:

1. Inorganic Chemistry ; Ajay kumar

2.Inorganic Chemistry By Madan Malik Tuli

3.principal of inoganic chemistry by puri sharma kalia

Reference Books:

1. Huheey, J. E.; Keiter, E.A. & Keiter, R.L. Inorganic Chemistry, Principles of Structure and Reactivity 4th Ed., Harper Collins 1993, Pearson,2006.
2. Sharpe, A.G. Inorganic Chemistry, 4th Indian Reprint (Pearson Education) 2005
3. Douglas, B. E.; McDaniel, D.H. & Alexander, J.J. Concepts and Models in Inorganic Chemistry3rd Ed., John Wiley and Sons, NY, 1994.
4. Greenwood, N.N. & Earnshaw, A. Chemistry of the Elements, Elsevier 2nd Ed, 1997 (Ziegler Natta Catalyst and Equilibria in Grignard Solution).
5. Lee, J.D. Concise Inorganic Chemistry 5th Ed., John Wiley and sons 2008.
6. Powell, P. Principles of Organometallic Chemistry, Chapman and Hall, 1988

Course outline

Nomenclature

Code	Course Title	T-P-Pj (Credit)	Prerequisite
CUTM1482	**Spectroscopy**	**3-2-1**	

1. Objective

- To learn proper sample handling procedures for acquiring infrared **spectra**.
- To determine the percentage composition of a liquid sample mixture by the application of Beer's law.
- To train the students to synthesize various dyes

2. Course outcome

COs	Course outcomes

CO1	To impart knowledge of basic concepts of spectroscopy
CO2	Understanding of the problem solving techniques of organic spectrosocpy
CO3	Investigate the structure of molecules on the basis of spectroscopic data
CO4	Investigation of different patterns of peaks and their origin and hence deduction of structure of organic species.
CO5	To develop analytical skills on predicting and explaining the various spectroscopic patterns for organic molecules

Course Outcome to Program Outcome Mapping:

COs	**PO1**	**PO2**	**PO3**	**PO4**	**PO5**	**PO6**	**PO7**	**PO8**	**PO9**	**PO10**	**PO11**	**PO12**	**PSO1**	**PSO2**	**PSO3**
CO1	3	2	2	3	3	1	3	2				3	3	2	3
CO2	3	3	3	3	3	1	2	1				3	3	2	3
CO3	-	3	2	3	3	1	2	1				2	3	2	3
CO4	2	3	3	3	3	2	2	2				3	3	2	3
CO5	2	2	2	3	3	2	3	2					3	2	3

***High-3, Medium-2, Low-1**

Course Content

Module -1(UV-VIS Spectroscopy) 6 hrs

Introduction to UV spectroscopy, Theory and principles, Types of electronic transitions, Tmax, Chromophores and Auxochromes, Bathochromic and Hypsochromic shifts, Intensity of absorption; of Application Woodward rules for calculation of λmax for the following systems: α,β unsaturated aldehydes, ketones, carboxylic acids and esters; Conjugated dienes: distinction between cis and trans isomers.

Assignment-1: Types of electronic transitions, Tmax, Chromophores and Auxochromes, Bathochromic and Hypsochromic shifts, Intensity of absorption
Practise-1: To study the UV-Visible Spectroscopy of Conjugated Molecules
Practise-2: Solvent Effects on the UV-visible Absorption Spectra
Practise-3: To Determine of the Molar Absorptivity of a Light Absorbing Molecule by using Lambert-Beer's law

Module-2 (IR Spectroscopy) 6hrs

Fundamental and non-fundamental molecular vibrations; IR absorption positions of O, N and S containing functional groups; Effect of H"bonding, conjugation, resonance and ring size on IR absorptions; Fingerprint region and its significance; application in functional group analysis.

Assignment-2:IR absorption positions of O, N and S containing functional groups
Assignment-3: Fingerprint region and its significance

Module -3 (NMR Spectroscopy) 6 hrs

Basic principles of Proton Magnetic Resonance, chemical shift and factors influencing it; Spin - Spin coupling and coupling constant; Anisotropic effects in alkene, alkyne, aldehydes and aromatics, Interpretation of NMR spectra of simple compounds. Applications of IR, UV and NMR for identification of simple organic molecules.

Assignment-4:chemical shift and factors influencing it
Assignment-5: Spin - Spin coupling and coupling constant

Module-4(Mass spectroscopy) 5 hrs

Introduction to mass spectrometry, Ionization methods & mass filter techniques in mass spectrometry, ionization methods(electron impact, chemical ionization), molecule fragmentation, Mac-Lafferty rearrangement

Assignment-6: Mac-Lafferty rearrangement

Module-5(Carbohydrates) 7 hrs

Occurrence, classification and their biological importance. Monosaccharides: Constitution and absolute configuration of glucose and fructose, epimers and anomers, mutarotation, determination of ring size of glucose and fructose, Haworth projections and conformational structures; Interconversions of aldoses and ketoses; Killiani-Fischer synthesis and Ruff degradation; Disaccharides - Structure elucidation of maltose, lactose and sucrose.Polysaccharides: starch, cellulose and glycogen.

Assignment-7: epimers and anomers
Assignment-8:mutarotation

Module – 6 (Dyes) **6hrs**

Classification, Colour and constitution; Mordant and Vat Dyes; Chemistry of dyeing; Synthesis and applications of: Azo dyes - Methyl Orange and Congo Red (mechanism of Diazo Coupling); Triphenyl Methane Dyes -Malachite Green, Rosaniline and Crystal Violet; Phthalein Dyes - Phenolphthalein and Fluorescein.
Assignment-9: Mordant and Vat Dyes
Practise-4:synthesis of fluorescein dye from marker ink and phthalic anhydride
Practise-5: Methyl Orange : Organic synthesis

Module-7 **(Chromatography)** **4hrs**
Gas Chromatography, Column Chromatography, supercritical fluids, TLC,HPLC

Assignment-10:TLC and supercritical fluids
Practise-6: Separation of compounds using column chromatography

Reference Books:

1. Organic spectroscopy by William Kemp
2. Morrison, R. T. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India)Pvt. Ltd. (Pearson Education
3. Spectrophotometric identification of Organic Compounds by Silverstein
4. Billmeyer, F. W. Textbook of Polymer Science, John Wiley & Sons, Inc.4. Gowariker, V. R.; Viswanathan, N. V. & Sreedhar, J. Polymer Science, New Age International (P) Ltd. Pub.
5. 4. Finar, I. L. Organic Chemistry (Volume 2: Stereochemistry and the Chemistry of Natural Products), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
6. 5. Clayden, J.; Greeves, N.; Warren, S.; Wothers, P.; Organic Chemistry, Oxford University Press.
7. Pavia, D. L. et al. Introduction to Spectroscopy 5th Ed. Cengage Course India Ed. (2015).